

FACTORS AFFECTING SOLUBILITY

Solubility

- We know from the **particle theory** that everything is made up of particles that are attracted to each other and are in constant motion
- In a solution, the solute and solvent particles are constantly bumping into each other, creating spaces and mixing until the different types of particles are distributed throughout the solution

Rate of Dissolving

- The **rate of dissolving** is how quickly a solute dissolves in a solvent
- This can be affected by three factors of solubility:

1) Stirring

2) Temperature

3) Particle Size

Factors Affecting Solubility

1) Stirring

- Stirring **increases** the rate at which a solute dissolves because **it pushes the particles of the solute away from each other, making space for the solvent particles**

2) Temperature

- Temperature **increases** the rate at which a solute dissolves because **at higher temperatures particles move faster, allowing the mixing process to be quick**

3) Particle Size

- Large particles **decreases** the rate at which a solute dissolves because **it is harder for the particles of solute and solvent to interact**