Name: _____

OUR TEST IS ON FRIDAY, APRIL 11, 2014

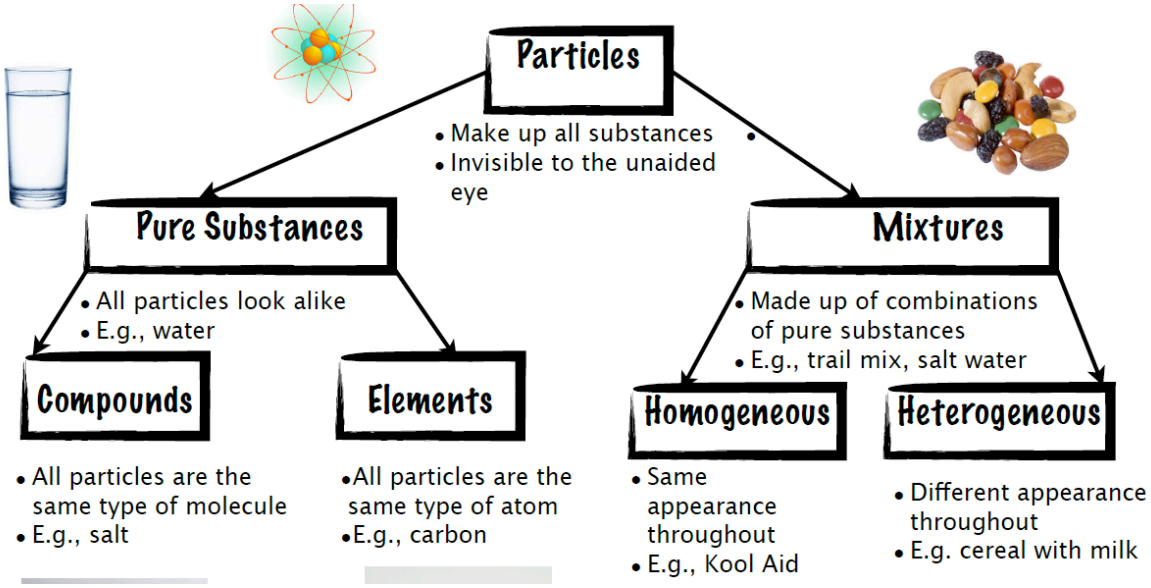
Grade Seven Science Test Review: Unit Three
Pure Substances and Mixtures

THINGS TO KNOW

Lesson 1: Matter

- ☐ *Matter* is anything that has mass and takes up space
- ☐ Examples of matter

Lesson 2-4: Pure Substances and Mixtures



- ☐ *Homogeneous* = SOLUTIONS (same appearance throughout)
- ☐ *Heterogeneous* = MECHANICAL MIXTURES (different appearance throughout)

Lesson 5-6: Particle Theory

- ☐ The six parts of Particle Theory
- ☐ Using Particle Theory to explain experiments

Lesson 7: Solutions

- ☐ *Solutions* are made of a *solute* (less of it) and a *solvent* (more of it)
- ☐ Solutions can be homogeneous or heterogeneous
- ☐ Solutes can be *soluble* or *insoluble* in any given solvent at any given temperature
- ☐ Solutions can be *concentrated* or *dilute*
- ☐ Solutions can be *undersaturated, saturated and supersaturated*

Lesson 8: Factors of Solubility

- ☐ Temperature
- ☐ Particle Size
- ☐ Stirring

Lesson 10-11: Separating Solutions and Mixtures

- ☐ Ways of separating solutions (chromatography,
- ☐ Ways of separating mechanical mixtures (magnetism, filtration, sorting, sifting)

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TYPES OF QUESTIONS

- Matching definitions
- Multiple choice
- True/ False, explain
- Diagrams
- Short answer

SAMPLE QUESTIONS

1. Use the terms “dilute” and “concentrated” to explain the difference between frozen juice concentrate and juice.

2. Explain how paper chromatography could be used to separate a mixture of different coloured inks.

3. List three factors that influence the rate at which dissolving occurs.

4. Use the particle theory to explain why hot chocolate powder dissolves more rapidly in hot water than cold water.

5. Explain the difference between an undersaturated, saturated and supersaturated solution.

6. Explain whether you would classify each of the following items as a pure substance or a mixture: apple juice, fruit punch, distilled water, lemonade.

7. Explain why a solid substance occupies less space than the same substance in its gaseous form.

HOW TO STUDY

- Review class notes and labs
- Revisit your science journals
- Review power points on class website (mskuksclass.weebly.com)

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